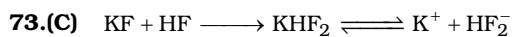




**72.** In case of  $(\text{SiH}_3)_3\text{N}$ , lone pair of electrons on nitrogen is involved in  $p\pi - d\pi$  back bonding, while in case of  $(\text{CH}_3)_3\text{N}$ , the  $p\pi - d\pi$  back bonding is not possible due to the absence of vacant d-orbitals in carbon. Because of this  $(\text{CH}_3)_3\text{N}$  is more basic than  $(\text{SiH}_3)_3\text{N}$ .



**74.(C)** We can represent  $\text{CsBr}_3$  as  $\text{Cs}^+\text{Br}_3^-$

